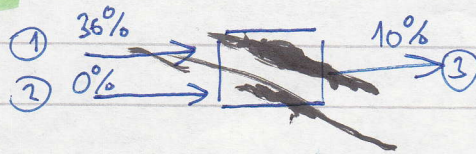


42.1

A = HCl

B = H₂O



w_{ij}	1	2	3
A - HCl	0.36	0	0.1
B - H ₂ O	0.64	1	0.9
$\sum m_j w_j$	59 kg	m_2	m_3

balance A

$$m_1 \cdot 0.36 = m_3 \cdot 0.1 \Rightarrow 59 \cdot 0.36 = m_3 \cdot 0.1$$

balance m

$$m_1 + m_2 = m_3 \Rightarrow 59 + m_2 = m_3$$

$$59 \cdot 0.36 = (59 + m_2) \cdot 0.1$$

$$59 \cdot 0.36 - 5.9 = 0.1 m_2$$

$$m_2 = 153.4 \text{ kg}$$

